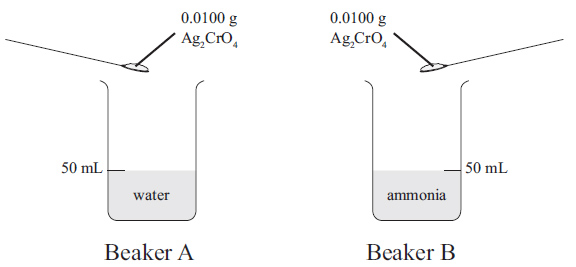
**Solubility of solids in solutions forming a complex ion**

**1)** In an experiment, 0.0100 g of Ag2CrO4 in beaker A was made up to a volume of 50.0 mL with

water. In beaker B, 0.0100 g of Ag2CrO4 was made up to a volume of 50.0 mL with 0.100 mol L–1 ammonia solution.



Compare and contrast the solubility of Ag2CrO4 in beaker A and beaker B. *No calculations are necessary.*

**2)** Solid sodium chloride is added to 5.00 L of 0.100 mol L–1 silver nitrate solution. Calculate the

minimum mass of sodium chloride that would be needed to produce a saturated solution of AgCl.

Assume that there is no change in volume when the sodium chloride is added. *M*(NaCl) = 58.5 g mol–1

Discuss reasons for the fact that a precipitate of silver chloride dissolves on the addition of excess aqueous

ammonia.

© 2015 <http://www.chemicalminds.wikispaces.com>

NCEA questions and answers reproduced with permission from NZQA